

1st law of thermodynamics simple

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There are four laws of thermodynamics. They talk about temperature, heat, work, and entropy. They are used in thermodynamics and other sciences, for example chemistry.

Thermodynamics has three main laws: the first law, the second law, and the third law. Then there was another law, called the "zeroth law." The law of conservation of mass is also an important idea in thermodynamics, but it is not called law.

Laws of Thermodynamics is an integral part of Chemistry, which not only explain the surrounding but also the infinite Universe which governs by the laws of chemistry and others.

The zeroth law of thermodynamics says temperature is an empirical parameter in thermodynamic systems. It states the transitive relationship between the temperatures of multiple bodies in thermal equilibrium. The law says:

If two systems are both in thermal equilibrium with a third system, then they are in thermal equilibrium with each other.

The first law of thermodynamics is a version of the law of conservation of energy, adapted for thermodynamic systems. The law of conservation of energy states that the total energy of an isolated system is constant; energy can be transformed from one form to another, but can be neither created nor destroyed. It can also be stated in the following form:

The energy gained (or lost) by a system is equal to the energy lost (or gained) by its surroundings.

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