Mercury cadmium vs alkaline batteries

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This is a list of commercially-available battery types summarizing some of their characteristics for ready comparison.

Types of dry-cell batteries are zinc-carbon batteries, alkaline-cell batteries, and mercury batteries. Before zinc-carbon batteries were used, mercury batteries were the main resource. It was not until mercury was known to become harmful that zinc-carbon batteries replaced it.

A different form of mercury battery uses mercuric oxide and cadmium. This has a much lower terminal voltage around 0.9 volts and so has lower energy density, but it has an extended temperature range, in special designs up to 180 C. Because cadmium has low solubility in the alkaline electrolyte, these batteries have long storage life. [4]

What if specific battery chemistries excel in some areas and are poor in others? In today"s post, we answer those questions by comparing six common battery chemistries" lifetime, cost, power/weight ratio, temperature range, storability and ease of disposal. Take this blog post with you!

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A mercury battery (also called mercuric oxide battery, mercury cell, button cell, or Ruben-Mallory[1]) is a non-rechargeable electrochemical battery, a primary cell. Mercury batteries use a reaction between mercuric oxide and zinc electrodes in an alkaline electrolyte. The voltage during discharge remains practically constant at 1.35 volts, and the capacity is much greater than that of a similarly sized zinc-carbon battery. Mercury batteries were used in the shape of button cells for watches, hearing aids, cameras and calculators, and in larger forms for other applications.

For a time during and after World War II, batteries made with mercury became a popular power source for portable electronic devices. Due to the content of toxic mercury and environmental concerns about its disposal, the sale of mercury batteries is now banned in many countries.[2] Both ANSI and IEC have withdrawn their standards for mercury batteries.

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